

## Chemistry B Moles Packet Answers

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# GRAMS Molar mass (grams per mole) # MOLES 6.02 X 10<sup>23</sup> atoms or molecules per mole # MOLECULES 22.4 liters of gas per mole # LITERS (at STP) EXAMPLE 1: What would be the volume in liters of 40.36 grams of neon at STP? # liters Ne = 40.36 g Ne x 1 mole Ne x 22.4 L Ne = 44.80 L Ne 20.18 g Ne 1 mole Ne EXAMPLE 2: How many molecules would there be in 56 liters of carbon dioxide at STP? # molecules CO<sub>2</sub> = 56.0 L CO<sub>2</sub> x 1 mole CO<sub>2</sub> x 6.02 X 10<sup>23</sup> molecules CO<sub>2</sub> = 1.51 x 10<sup>24</sup> molecules CO<sub>2</sub> 22.4 L CO<sub>2</sub> 1 ...

~~CHEMISTRY B- MOLES PACKET NAME: HR: \_\_\_\_\_ PAGE 1~~

CHEMISTRY B- MOLES PACKET NAME: \_\_\_\_\_ HR: \_\_\_\_\_ PAGE 3 CALCULATE THE MOLAR MASS FOR THE FOLLOWING COMPOUNDS OR DIATOMIC ELEMENTS. SET UP EACH PROBLEM AS SHOWN IN THE EXAMPLE ABOVE. INCLUDE UNITS (G/MOL) 1. water H 2O 2. calcium chloride CaCl<sub>2</sub> 3. copper(II) sulfate CuSO<sub>4</sub> 4. silver nitrate AgNO<sub>3</sub> 5. sulfuric acid H

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CHEMISTRY B- MOLES PACKET NAME: \_\_\_\_\_ HR: \_\_\_\_\_ PAGE 6CHEMISTRY WORKSHEET # 4 MOLAR MASSES AND AVOGADRO'S NUMBERNow that you know two definitions of a mole (a molar mass and an Avogadro's number of particles) you cancombine these two definitions into one problem.EXAMPLE: How many molecules are there in 90.1 grams of water?2 H = 2 x(1.01) = 2.02O = 1 x (16.00) = 16.00H 2 O= 18.02 g? molecules H 2 O = 90.1 g H 2 O x 1 mole H 2 O x 6.02 x 10<sup>23</sup> molecules H 2 O = 3.01 x 10<sup>24</sup> molecules H 2 ...

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